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Zinc is determined by EDTA titration in almost the same way magnesium is - in the pH10 against Eriochrome BlackT. If the solution initially contains also different metal ions, they should be removed or masked, as EDTA react easily with most cations (with the exception of alkali metals). reaction Reaction taking place during titration is Zn2+ + EDTA4- ZnEDTA2- sample size For 0.01M titrant and assuming 50mL burette, aliquot taken for titration should contain about 0.35-0.45 millimoles of zinc (23-29mg). If preparation of such sample is difficult, we can use different EDTA concentration. end point detection End point of zinc titration is easily detected with Eriochrome BlackT. solutions used To perform titration we will need titrant - 0.01M EDTA solution and ammonia pH10.0 buffer. We will also need indicator - either in the form of solution, or ground with NaCl - 100mg of indicator plus 20g of analytical grade NaCl. procedure Procedure to follow doesn't differ from the one used for the magnesium titration. Transfer zinc solution to Erlenmeyer flask. Dilute to about 100mL with distilled water. Add 2mL of pH10 ammonia buffer solution. Add a pinch of Eriochrome BlackT ground with sodium chloride (100mg of indicator plus 20g of analytical grade NaCl). Titrate with EDTA solution till the color changes to blue. result calculation Calculation of EDTA titration results is always easy, as EDTA reacts with all metal ions in 1:1 ratio: Zn2+ + EDTA4- ZnEDTA2- That means number of moles of zinc is exactly that of number of moles of EDTA used. To calculate zinc solution concentration use EBAS - stoichiometry calculator. Download determination of zinc reaction file, open it with the free trial version of the stoichiometry calculator.Click n=CV button above EDTA4+ in the input frame, enter volume and concentration of the titrant used. Click Use button. Read mass of zinc in the titrated sample in the output frame. In general this is a simple titration, with no other problems then those listed as general sources of titration errors. Zinc is determined by EDTA titration in almost the same way magnesium is - in the pH10 against Eriochrome BlackT. If the solution initially contains also different metal ions, they should be removed or masked, as EDTA react easily with most cations (with the exception of alkali metals). reaction Reaction taking place during titration is Zn2+ + EDTA4- ZnEDTA2- sample size For 0.01M titrant and assuming 50mL burette, aliquot taken for titration should contain about 0.35-0.45 millimoles of zinc (23-29mg). If preparation of such sample is difficult, we can use different EDTA concentration. end point detection End point of zinc titration is easily detected with Eriochrome BlackT. solutions used To perform titration we will need titrant - 0.01M EDTA solution and ammonia pH10.0 buffer. We will also need indicator - either in the form of solution, or ground with NaCl - 100mg of indicator plus 20g of analytical grade NaCl. procedure Procedure to follow doesn't differ from the one used for the magnesium titration. Transfer zinc solution to Erlenmeyer flask. Dilute to about 100mL with distilled water. Add 2mL of pH10 ammonia buffer solution. Add a pinch of Eriochrome BlackT ground with sodium chloride (100mg of indicator plus 20g of analytical grade NaCl). Titrate with EDTA solution till the color changes to blue. result calculation Calculation of EDTA titration results is always easy, as EDTA reacts with all metal ions in 1:1 ratio: Zn2+ + EDTA4- ZnEDTA2- That means number of moles of zinc is exactly that of number of moles of EDTA used. To calculate zinc solution concentration use EBAS - stoichiometry calculator. Download determination of zinc reaction file, open it with the free trial version of the stoichiometry calculator.Click n=CV button above EDTA4+ in the input frame, enter volume and concentration of the titrant used. Click Use button. Read mass of zinc in the titrated sample in the output frame. In general this is a simple titration, with no other problems then those listed as general sources of titration errors.

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